

**MCQ with one correct answer**

1. The rate of reaction that does not involve gases, is not depend on [     ]
  - a) Pressure
  - b) Temperature
  - c) Concentration
  - d) catalyst
2. In general, with every 10<sup>0</sup> C rise in temperature, the rate of reaction becomes approximately [     ]
  - a) Ten times
  - b) Doubled
  - c) Halved
  - d) One tenth
3. A positive Catalyst is used to [     ]
  - a) Increases the product
  - b) Decreases the product
  - c) Minimise the time of reaction
  - d) None of the above
4. Concentration of reactants increases then rate of reaction [     ]
  - a) Decreases
  - b) Increases
  - c) Does not change
  - d) Zero
5. Among the following the incorrect statement is [     ]
  - a) Rate of reaction is directly proportional to concentration of reactants
  - b) Rate of reaction is directly proportional to temperature
  - c) Rate of reaction is directly proportional to particle size
  - d) Rate of reaction is directly proportional to surface area

**CLASS ROOM ASSIGNMENT**

**ORANGE WORKSHEET**

- I. **MCQ with one correct answer :**
  1. A catalyst accelerates the reaction because [     ]
    - a) It brings the molecules closer
    - b) It lowers the activation energy
    - c) It changes the rate of reaction
    - d) It increases the activation energy
  2. The factor(s) which influence the rate of reaction is / are [     ]
    - a) Nature of reactant
    - b) Concentration of the reactants
    - c) Temperature
    - d) All the three

**II. Statement type :**

3. **Statement (A) :** Rate of a reaction is usually doubled (or) tripled as temperature is increased by 10<sup>0</sup>C.  
**Statement (B) :** Rate of photo chemical reactions also depend upon the intensity of light exposed to reaction mixture. [     ]  
**Statement (C) :** A catalyst may increase (positive catalyst) or decrease (negative catalyst) the rate of a reaction without undergoing any chemical change  
  - a) All the above statements are correct
  - b) All the above statements are incorrect
  - c) A, B are correct and C is incorrect
  - d) A, B are incorrect and C is correct.

**YELLOW WORKSHEET**

**I. MCQ with one correct answer :**

4. An elementary reaction is given as  $2P + Q \rightarrow \text{products}$ . If concentration of Q is kept constant and concentration of P is doubled the rate of reaction is [     ]
  - a) Doubled
  - b) Halved
  - c) Quardrupled
  - d) Remains same